

Name: \_\_\_\_\_

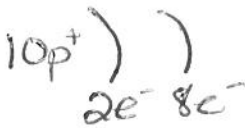
Date: \_\_\_\_\_

### The Bohr-Rutherford Model

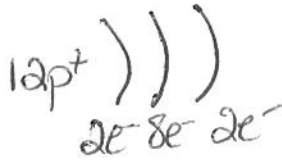
Draw the Bohr-Rutherford model for each of the following atoms.

1. Write down the number of electrons (e-), protons (p+), and neutrons (n°) for each of these atoms.
2. How many valence electrons should each atom have in its outer shell?
3. Draw the atoms using the Bohr-Rutherford Model.

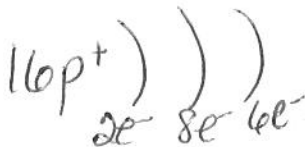
Ne  
2 orbits  
e- = 10  
P+ = 10  
n° = 10  
Valence = 8



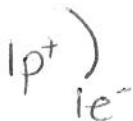
Mg  
3 orbits  
e- = 12  
P+ = 12  
n° = 12  
Valence = 2



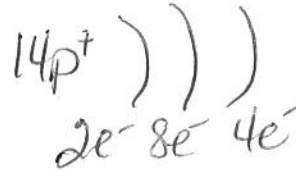
S  
3 orbits  
e- = 16  
P+ = 16  
n° = 16  
Valence = 6



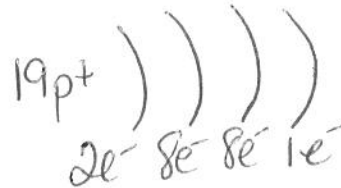
H  
1 orbit  
e- = 1  
P+ = 1  
n° = 0  
Valence = 1



Si  
3 orbits  
e- = 14  
P+ = 14  
n° = 14  
Valence = 4



K  
4 orbits  
e- = 19  
P+ = 19  
n° = 20  
Valence = 1



C  
2 orbits  
e- = 6  
P+ = 6  
n° = 6  
Valence = 4

